Tiara-like Octanuclear Palladium(II) and Platinum(II) Thiolates and Their Inclusion Complexes with Dihalo- or Iodoalkanes

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Abstract
A tiara-like octanuclear palladium thiolate complex, [Pd(μ-SCH₂CO₂Me)₂]₈ that has a toroidal structure, was synthesized via reactions of either PdCl₂ with methyl thioglycolate/N,N-diisopropylethylamine (DIEA) (conventional method) or [PdCl₂(MeCN)₂] with m-C₆H₄(CMe₂SCH₂CO₂Me)₂ (alternative method). In the latter method, the tiara-like complex formed via the corresponding SCS-pincer complex and/or 1:1 PdCl₂ and ligand complexes. With respect to the platinum analogs, the alternative method efficiently produced the tiara-like octanuclear complex, [Pt(μ-SCH₂CO₂Me)₂]₈, in high purity. Small molecules, such as CH₂Cl₂, ClCH₂CH₂Cl, CH₂Br₂, and CH₃I, were accommodated in the inner voids of the tiara rings to form 1:1 inclusion complexes. These complexes are stabilized by weak CH⋯X hydrogen bonds (X = Cl or Br) between the methylene protons of four or eight axially positioned methoxycarbonylmethyl groups on the tiara rings and the halogen atoms of the guest molecules, as well as weak coordination of the halogen atoms to the transition-metal atoms.
Introduction

Metal thiolates are an important class of compounds because of their structural diversity in coordination chemistry,^{1-3} relevance to cofactors of metalloproteins in biological systems,^{4-5} and importance in cluster and surface science^{6-7} and catalysis.^{8} Thiolate-bridged polynuclear group 10 transition-metal complexes with chain structures, \([M(\mu-SR)]_n \) (\(M = \text{Ni, Pd}\)), have been investigated in detail because they provide efficient catalysis of atom-economical organic reactions such as regioselective additions of thiols and disulfides across alkynes.\(^{8-13}\) Tiara-like complexes, which are also polynuclear group 10 transition-metal thiolates and are characterized by toroidal architectures, have been studied extensively, both in terms of their intriguing structural features and with respect to the preparation of monodisperse metal sulfide nanoparticles,\(^{14}\) non-linear optical materials,\(^{15}\) photoactive water-reducing catalysts,\(^{16-17}\) and host-guest chemistry.\(^{15, 18-21}\) Tiara-like nickel complexes have received considerable attention resulting in the preparation of complexes with a variety of ring sizes \([\text{Ni}(\mu-SR)]_n \) \((n = 4-6, 8-12)\).\(^{15, 17-18, 20-35}\) In contrast, although several tiara-like hexanuclear palladium complexes have been reported,\(^{9, 14, 19, 36-43}\) only one octanuclear complex, i.e., \([\text{Pd}(\mu-S^8\text{Pr})_2]_6\), is known, and was obtained as a mixture with hexanuclear \([\text{Pd}(\mu-S^8\text{Pr})_2]_6\).\(^{19}\) For platinum, infinite thiolates and selenolates \([\text{Pt}(\mu-\text{E})_2]_\infty \) \((E = S, \text{Se})\) have been synthesized using a solvothermal method;\(^{37}\) however, to the best of our knowledge, no tiara-like complexes have been reported thus far.

In most cases, tiara-like nickel and palladium complexes are synthesized through the reactions of group 10 transition-metal halides with thiols in the presence or absence of a base. We successfully synthesized an octanuclear palladium complex using this conventional method. Moreover, we also developed an alternative synthetic method using \(m\)-C\(_6\)H\(_4\)(CMe\(_2\)SR)\(_2\) as the thiolate ligand source. The latter method was found to be advantageous for the synthesis of a platinum complex. In this paper, we report the synthesis of novel tiara-like octanuclear palladium and platinum complexes using these two methods, and their inclusion complexes containing small organic molecules. Decanuclear and undecanuclear nickel complexes have been reported to include solvent molecules, such as benzene,\(^{20}\) toluene,\(^{15}\) and THF,\(^{21}\) in their inner voids. In contrast, no tiara-like palladium inclusion complexes have been reported thus far.

Results and Discussion

Synthesis and Structure of Tiara-like Octanuclear Palladium Complex 1

\(\text{PdCl}_2\) was reacted with two equivalents each of methyl thioglycolate and DIEA in \(n\)-PrOH at room temperature. After 4 h, a tiara-like octanuclear complex, i.e., \([\text{Pd}(\mu-\text{SCH}_2\text{CO}_2\text{Me})_2]_8\) \((1)\) was obtained in high isolated yield \((90\%)\) (Scheme 1); a high concentration of palladium \((0.25 \text{ M})\) is essential to achieve a high yield because a lower concentration \((0.04 \text{ M})\) resulted in lower yield \((67\%)\). The molecular structure of 1 was determined by X-ray crystallography (Figure 1, recrystallized from CHCl\(_3\)/hexane). The eight palladium atoms lie on a nearly planar surface and form an octagonal structure with the
distances between two adjacent metal atoms ranging from 3.1398(9) to 3.2745(10) Å; the average Pd····Pd distance is 3.23(4) Å, which is slightly longer than the metal–metal distances in [Pd(μ-S^Pr)]_8 (3.17 Å)\(^{19}\) and an octanuclear nickel analog, [Ni(μ-SCH_2CO_2Et)]_8 (3.05 Å).\(^{18}\) Each palladium atom is coordinated by four μ_2-bridged thiolate ligands and has an almost square planar geometry. The Pd–S bond distances are within 2.305(2)–2.367(4) Å (average: 2.32(2) Å). The horizontal and vertical S–Pd–S bond angles vary between 93.50(14) and 99.21(8)° (average: 97.3(14)°), and 81.65(8) and 85.35(14)° (average: 83.0(9)°), respectively. The Pd–S–Pd bond angles are within 85.20(8)–91.70(17)° (average: 88.4(16)°), and the dihedral angles between the PdS_4 planes are within 127.75(6)–138.41(6)° (average: 135(4)°). The distances between two adjacent sulfur atoms in the two S_8 rings are within the range of 3.336(6)–3.573(6) Å (average: 3.48(5) Å). The sixteen methoxycarbonylmethyl arms are alternately located at axial and equatorial positions (i.e., the arms extend nearly perpendicularly and horizontally from the tiara ring), and some are significantly disordered; in particular, the axial arm on the S1 atom alternates between bending inward to occupy the inner void and extending out of the toroidal structure (~1:1). A similar inward curvature of an arm was also observed in [Ni(μ-SCH_2CO_2Et)]_8.\(^{18}\) The axial arm on the S2 atom, which is located trans to S1 on the other rim of the tiara, closes the lid from the bottom. The distances between the diagonally positioned palladium atoms are Pd1····Pd5 = 8.7191(9) Å, Pd2····Pd6 = 8.2331(9) Å, Pd3····Pd7 = 8.2307(8) Å, and Pd4····Pd8 = 8.5895(8) Å, resulting in an ellipsoidal architecture of the tiara ring (average: 8.44 Å). These distances are considerably longer than those of [Ni(μ-SCH_2CO_2Et)]_8 (7.59–8.34 Å, average: 7.95 Å).\(^{18}\) The above-mentioned geometrical parameters for 1 were compared with those of the previously reported hexanuclear palladium complexes [Pd(μ-SCH_2CO_2Me)]_6\(^{39}\) and [Pd(μ-S^Hex)]_6.\(^9\) The average distance between the two adjacent palladium atoms in the hexanuclear complexes is 3.11 Å, which is approximately 0.1 Å shorter than that in 1. Although the averages of the horizontal and vertical S–Pd–S bond angles are similar for these three complexes (97.8(5)° and 82.4(5)° for [Pd(μ-SCH_2CO_2Me)]_6, and 98.0(5)° and 81.9(5)° for [Pd(μ-S^Hex)]_6, the average Pd–S–Pd bond angle for the hexanuclear complexes is 84.5°, which is approximately 4° smaller than that found for 1. The dihedral angles between the PdS_4 planes are also small: 113.47(3)–124.36(4)° for [Pd(μ-SCH_2CO_2Me)]_6 and 115.19(6)–126.53(7)° for [Pd(μ-S^Hex)]_6. The distances between the diagonally positioned palladium atoms in [Pd(μ-S^Hex)]_6 are within 6.053–6.461 Å. The average distance (6.23 Å) is approximately 2.2 Å shorter than that in 1.
**Scheme 1.** Formation of Tiara-like Octanuclear Palladium Complex 1

\[
\text{HS} \xrightarrow{\text{CO}_2\text{Me} \text{ (2.0 equiv)}} \text{DIEA (2.0 equiv)}} \xrightarrow{\text{n-PrOH, r.t., 4 h}} \]

\[
1: 90\% (R = \text{CH}_3\text{CO}_2\text{Me})
\]

**Figure 1.** Molecular structure of 1 with an inwardly bent arm (left) and outwardly stretched arm (right) on S1 atom (upper: top view, lower: side view). Hydrogen atoms are omitted for clarity. Pd: light green, S: orange, C: gray, O: red.

The MALDI-TOF MS spectrum of 1 revealed a parent peak at \( m/z \) 2556.3 (with \( \text{Na}^+ \), calcd 2556.2) and
a small fragment peak at $m/z$ 2240.4 assignable to \([\text{Pd(µ-SCH}_2\text{CO}_2\text{Me})_2]_7\) (with Na\(^+\), calcd 2238.3) (Figures S1–S3 in Supporting Information). The \(^1\)H NMR spectrum of 1 (in CDCl\(_3\)) features two methylene signals at 3.34 and 3.23 ppm and two methoxy signals at 3.97 and 3.77 ppm; these two sets of signals are attributed to the axial and equatorial methoxycarbonylmethyl arms, and therefore, the molecular symmetry of 1 in solution is pseudo-$D_{4d}$. Such separated NMR signals of axial and equatorial arms have also been observed in hexanuclear palladium complexes.\(^9\), \(^14\), \(^19\), \(^41\)-\(^42\) In contrast with \([\text{Pd(µ-S}^\text{S}^\text{Pr})_2]_8\), which exists in equilibrium with the corresponding hexanuclear complex \([\text{Pd(µ-S}^\text{S}^\text{Pr})_2]_6\),\(^19\) 1 was stable in solution (CD\(_3\)CN) even at 70 °C, and no hexanuclear or other polynuclear complexes were observed at all. The \(^1\)H NMR spectra of 1 in CDCl\(_3\) and CD\(_2\)Cl\(_2\) obtained through low-temperature measurements are shown in Figure 2. In CDCl\(_3\), one of the two methoxy signals appears at 3.97 ppm as a slightly broad singlet, which is broadened further compared with the other signal at 3.77 ppm upon decreasing the temperature. On the other hand, in CD\(_2\)Cl\(_2\), both the methoxy signals broaden simultaneously. This difference can be attributed to the fact that CDCl\(_3\) cannot be accommodated in the inner void of 1, but CD\(_2\)Cl\(_2\) can (vide infra). Instead of the solvent molecule, one of the axial methoxycarbonylmethyl groups can occupy the inner void in CDCl\(_3\), as already shown in Figure 1. The selective peak broadening observed in the variable-temperature NMR spectra in CDCl\(_3\) would result from the dynamic behavior of the axial arms moving into and out of the tiara ring. Therefore, the singlet that appears at 3.97 ppm is assigned to the methoxy groups in the axial arms, and that at 3.77 ppm is attributed to the methoxy groups in the equatorial arms. Because carbonyl carbons (170.4 and 169.8 ppm) have correlations with the axial methoxy protons (3.97 ppm) and methylene protons (3.34 ppm), as well as the equatorial methoxy protons (3.77 ppm) and other methylene protons (3.23 ppm) in the HMBC spectrum measured in CDCl\(_3\), respectively, the singlet at 3.34 ppm can be assigned to the methylene protons in the axial arms and the other at 3.23 ppm to the methylene protons in the equatorial arms.
Figure 2. Variable-temperature $^1$H NMR spectra (400 MHz) of 1 (left: in CDCl$_3$, -55–20 °C, right: in CD$_2$Cl$_2$, -85–20 °C).

Alternative Synthesis of 1
Although the conventional method was successful, complex 1 could also be prepared using dimethyl [1,3-phenylenebis(1-methylethylidenethio)]diacetate (2) as a thiolate-ligand source in the absence of a base. [PdCl$_2$(MeCN)$_2$] was reacted with one equivalent of 2 in MeOH (0.3 M) at 70 °C (bath temperature). After 10 h, complex 1 was obtained in high yield (Scheme 2).

Scheme 2. Alternative Synthesis of 1

To investigate the reaction mechanism, [PdCl$_2$(MeCN)$_2$] was reacted with 2 in CHCl$_3$ (0.04 M) at room temperature. After 1 h, SCS-pincer complex 3 formed selectively with 83% isolated yield; 1 was not formed at this stage (Scheme 3). After isolation, 3 was heated in CD$_3$CN (0.3 M) at 90 °C (bath temperature) in the presence of one equivalent of HCl, resulting in the formation of 1 in moderate yield along with a small amount of 1,3-diisopropenylbenzene (4). These reactions clearly demonstrate that tiara-like complex 1 can form via SCS-pincer complex 3. In contrast, when dimethyl [1,3-phenylenebis(methylenethio)]diacetate (5), which does not contain methyl groups at the benzylic
positions, was used instead of 2, higher temperatures and longer reaction times were required to generate SCS-pincer complex 6 (Scheme 4, Figure 3).44,45 When the same reaction was performed at room temperature, it produced an insoluble solid in nearly quantitative yield; the product appeared to be a monomer and/or oligomers of a 1:1 complex of PdCl₂ and 5 according to the results of elemental analyses (Anal. Caled for C₁₄H₁₈Cl₂O₄PdS₂: C, 34.19; H, 3.69. Found: C, 34.11; H, 3.56.).45 Thus, the steric hindrance derived from the four methyl groups at the benzylic positions in 2 facilitates the aromatic C(sp²)–H bond activation by placing the palladium atom proximate to the C(sp²)–H bond at the 2-position. Even at 90 °C, 1 did not form from the mixtures of either [PdCl₂(MeCN)₂] and 5 or 6 and HCl (Scheme 4).

**Scheme 3. Formation of Tiara-like Complex 1 via SCS-Pincer Complex 3**

```
[\text{PdCl}_2(\text{MeCN})_2] \rightarrow \begin{array}{c}
\text{2 (1.0 equiv)} \\
\text{HCl/MeCN (1.0 equiv)} \\
\text{MeCN} \\
\end{array} \\
\text{r.t., 1 h} \\
\text{90 °C (bath temp.), 10 h} \\
\rightarrow \begin{array}{c}
\text{3: 83%} \\
\text{4: 6%} \\
\end{array}
```

**Scheme 4. Formation and Reactivity of SCS-Pincer Complex 6**

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[\text{PdCl}_2(\text{MeCN})_2] \rightarrow \begin{array}{c}
\text{5 (1.0 equiv)} \\
\text{HCl/MeCN (1.0 equiv)} \\
\text{MeCN} \\
\end{array} \\
\text{r.t., 24 h} \\
\text{57% (90 °C, 20 h)}
```
Figure 3. Molecular structure of 6. Thermal ellipsoids are shown at a 50% probability level. Hydrogen atoms are omitted for clarity. Selected bond distances (Å), bond angles (°), and dihedral angles (°):
Pd1–C1 = 1.980(2), Pd1–S1 = 2.2967(7), Pd1–S2 = 2.3151(7), Pd1–Cl1 = 2.3961(7), C1–Pd1–S1 = 83.11(6), C1–Pd1–S2 = 84.46(6), S1–Pd1–Cl1 = 96.27(3), S2–Pd1–Cl1 = 96.12(2), Pd1–S1–C2 = 99.24(7), Pd1–S1–C4 = 105.61(7), C2–S1–C4 = 104.64(10), Pd1–S2–C3 = 97.67(7), Pd1–S2–C5 = 102.18(7), C3–S2–C5 = 100.81(11), C1–Pd1–S1–C2 = 25.12(10), C1–Pd1–S2–C3 = 24.22(10), C1–Pd1–S1–C4 = 133.26(10), C1–Pd1–S2–C5 = 78.64(10).

A possible mechanism for the formation of complex 1 from [PdCl\(_2\)(MeCN)\(_2\)] and 2 is shown in Scheme 5. After the generation of SCS-pincer complex 3 via aromatic C(sp\(^2\))–H bond activation, one of the thiolate groups is eliminated to provide a tertiary carbocation at the benzylic position with concurrent dissociation of a chloride ligand from palladium. This step rationalizes the difference in the reactivity of 3 and 6 because the generation of a carbocation intermediate from 6 is less favorable. Subsequent deprotonation affords a carbon–carbon double bond. The protonolysis of the aromatic C(sp\(^2\))–Pd bond by HCl followed by E1 elimination similar to that described above occurs again to give 4 and a mononuclear palladium thiolate, [Pd(SCH\(_2\)CO\(_2\)Me)\(_2\)], which then aggregates to construct a stable tiara-like octanuclear architecture. Although the above mechanism explains the formation of 1, the bidentate coordination of 2 to PdCl\(_2\) without aromatic C(sp\(^2\))–H bond activation\(^{45}\) and the subsequent stepwise elimination of the thiolate groups, similar to the above mechanism, can proceed concurrently to form 1.
Scheme 5. Possible Mechanism for the Formation of 1 from [PdCl₂(MeCN)₂] and 2

For examining whether a small amount of the hexanuclear complex [Pd(µ-SCH₂CO₂Me)₂]₆ is formed through the conventional method (Scheme 1) or the alternative method (Scheme 2), the crude reaction mixtures were analyzed by MALDI-TOF MS. No peak assignable to [Pd(µ-SCH₂CO₂Me)₂]₆ (with Na⁺, calcd 1921.4) was detected. Both these methods afforded 1 selectively; therefore, the predominant formation of the octanuclear complex over a hexanuclear complex seems not from the reaction conditions, but from the structure of the substituents on the sulfur atoms, i.e., the methoxycarbonylmethyl groups, which can be regarded as a template. In the case of nickel, the aforementioned [Ni(µ-SCH₂CO₂Et)₂]₈ is the sole example of an octanuclear tiara-like complex, reported by Dance et al. They mentioned that the size of the tiara ring can be determined by the van der Waals volume of the ethoxycarbonylmethyl group, which, similar to that of 1, also bends inward to occupy the inner void in the crystal structure. Dahl et al. proposed that macrocyclic [Ni(µ-SPh)₂]₁₁ was constructed around a THF molecule, which operated as a template during the growth of a ribbon-like nickel thiolate oligomer. On the other hand, [Pd(µ-SCH₂CO₂Me)₂]₆, which is a hexanuclear analog of 1, was prepared from PdCl₂⁻ and thioglycolic acid in MeOH. In this reaction, we presume that the formation of the hexanuclear tiara ring with carboxylic acids as substituents occurred first, and ester condensation between the carboxylic acid moieties and MeOH followed.

Synthesis and Structure of Tiara-like Octanuclear Platinum Complex 7
The synthesis of tiara-like platinum complexes was first examined using PtCl₂/methyl thioglycolate/DIEA similar to the reaction conditions described in Scheme 1 (in MeOH, 0.4 M, 80 °C, 10 h). A tiara-like octanuclear complex, [Pt(µ-SCH₂CO₂Me)₂]₈ (7), was formed in moderate yield (~60%); however, by-products, which may contain tiara-like complexes with different-sized rings as speculated from the ¹H NMR analysis, were formed concurrently, and the isolation of 7 from the mixture was unsuccessful. MALDI-TOF MS analysis of the crude reaction mixture showed small peaks assignable to [Pt(µ-SCH₂CO₂Me)₂]₀ at m/z 3673.3 (with Na⁺, calcd 3671.2) and [Pt(µ-SCH₂CO₂Me)₂]₁₀.
at m/z 4075.4 (with Na⁺, calcd 4076.5) in addition to a parent peak assignable to 7 at m/z 3267.7 (with Na⁺, calcd 3264.7); no peaks corresponding to tiara-like complexes with smaller rings were detected. In contrast, when PtCl₂ was reacted with one equivalent of 2 in MeCN (1 M) at 90 °C (bath temperature), 7 was obtained in high isolated yield with high purity (Scheme 6). The molecular structure of 7 was also determined by X-ray crystallography (Figure 4, recrystallized from CHCl₃/octane). In contrast to 1, the inside of the ring was void and a solvent molecule (CHCl₃) was located outside the ring. The range of the distances between two adjacent platinum atoms is 3.2409(4)–3.3345(4) Å (average: 3.29(3) Å); these distances are slightly longer than those between two adjacent palladium atoms in 1. The Pt–S bond distances range within 2.298(4)–2.356(6) Å (average: 2.32(1) Å) and are almost same as the Pd–S bond distances in 1. The horizontal and vertical S–Pt–S bond angles vary within 94.88(19)–101.82(15)° (average: 98.4(16)°) and 79.42(14)–85.3(2)° (average: 81.8(12)°), respectively, while the Pt–S–Pt bond angles vary within 88.2(2)–92.3(7)° (average: 90.3(12)°). According to the comparison of the S–M–S and M–S–M bond angles in 1 and 7 (M = Pd and Pt, respectively), the tiara ring of 7 is more horizontally elongated than that of 1 as a result of the changes in these angles. The dihedral angles between the PtS₄ planes range within 129.93(13)–142.25(8)° (average: 135(4)°). The distances between two adjacent sulfur atoms in the two S₈ rings are within the range of 3.409(8)–3.615(6) Å (average: 3.51(4) Å) and are also slightly longer than those in 1. The distances between the diagonally positioned platinum atoms are Pt₁···Pt₅ = 8.8011(5) Å, Pt₂···Pt₆ = 8.7539(5) Å, Pt₃···Pt₇ = 8.3559(5) Å, and Pt₄···Pt₈ = 8.4346(5) Å, resulting in an ellipsoidal architecture; the average distance (8.59 Å) is longer than that in 1 by ~0.15 Å.

Scheme 6. Formation of Tiara-like Octanuclear Platinum Complex 7
Figure 4. Molecular structure of 7 (upper: top view, lower: side view). A solvent molecule (CHCl₃) outside of the tiara ring and hydrogen atoms are omitted for clarity. Pt: pink, S: orange, C: gray, O: red.

The MALDI-TOF MS spectrum of isolated 7 reveals a parent peak at m/z 3265.7 as well as peaks at m/z 3191.7 and 3118.6 (Figures S4–S7 in Supporting Information), which are assigned to 7 with one or two methoxycarbonylmethyl group(s) missing, respectively (with Na⁺, calcd 3190.7 and 3117.7). The ¹H NMR spectrum of 7 (in CDCl₃) reveals two methylene signals at 3.53 and 3.42 ppm and two methoxy signals at 3.96 and 3.76 ppm, which appeared in the same fashion as those in 1. The variable-temperature measurements of 7 in CDCl₃ and CD₂Cl₂ indicated that the dynamic behavior of the axial arms was also similar to that observed in 1 (Figure 5), whereas no inward bend of the arms was observed in the crystal structure of 7. Compared with 1, the axial methoxy signal at 3.96 ppm was more broadened at low temperatures in CDCl₃. The HMBC spectrum of 7 measured in CDCl₃ also revealed correlations between carbonyl carbons (169.9 and 170.1 ppm) and the axial methoxy protons (3.96 ppm).
and methylene protons (3.53 ppm), as well as the equatorial methoxy protons (3.76 ppm) and other methylene protons (3.42 ppm), respectively; therefore, the singlet at 3.53 ppm is assignable to the methylene protons in the axial arms and the other singlet at 3.42 ppm to the methylene protons in the equatorial arms. These observations indicate that the methoxycarbonylmethyl groups should operate as a template during the formation of 7 in a manner similar to that of 1.

Figure 5. Variable-temperature $^1$H NMR spectra (400 MHz) of 7 (left: in CDCl$_3$, -55–20 °C, right: in CD$_2$Cl$_2$, -85–20 °C).

A more dilute reaction (0.04 M) resulted in the formation of SCS-pincer complex 8 (15%) in addition to 7 (57%). After isolation, 8 was heated to 90 °C in CD$_3$CN (0.4 M) in the presence of one equivalent of HCl; however, 7 did not form, thereby this reaction differs from the corresponding reaction involving palladium (Scheme 7). These results indicate that the formation of 7 does not proceed via SCS-pincer complex 8 but rather through monomeric and/or oligomeric 1:1 complexes of PtCl$_2$ and 2. The plausible formation mechanism of 7 from the 1:1 complexes is analogous to the mechanism described for the palladium complex above.

Scheme 7. Reaction of SCS-pincer Complex 8 with HCl

A trial for the formation of tiara-like palladium/platinum mixed complexes was performed through the reaction of 1 with 7 in CD$_3$CN at 70 °C. However, these complexes remained intact even after 48 h, and
mixed complexes were not obtained, as confirmed by the $^1$H NMR and MALDI-TOF MS measurements.

**Inclusion Complexes with Dihalo- or Iodoalkanes**

The diameters of the inner voids of 1 and 7 are ~5.1–5.2 Å, taking into account the van der Waals radii of palladium (1.63 Å) and platinum (1.75 Å). We found that the recrystallization of 1 and 7 from the mixtures of good solvents, such as CH$_2$Cl$_2$, ClCH$_2$CH$_2$Cl, CH$_2$Br$_2$, and CH$_3$I, and poor solvents, such as hexane and Et$_2$O gave the corresponding inclusion complexes, i.e., 1·CH$_2$Cl$_2$, 1·ClCH$_2$CH$_2$Cl, 1·CH$_2$Br$_2$, 1·CH$_3$I, 7·CH$_2$Cl$_2$, 7·ClCH$_2$CH$_2$Cl, and, 7·CH$_2$Br$_2$, respectively. The top and side views of the molecular structures of these complexes are shown in Figures 6 and 7. The guest molecules are disordered in these inclusion complexes, except for in 1·CH$_2$Cl$_2$; only the configurations with highest occupancy are depicted. In 1·CH$_2$Cl$_2$ and 1·CH$_2$Br$_2$, one of the two C–X bonds (X = Cl, Br) is arranged vertically and the other C–X bond is nearly horizontal. In contrast, in 7·CH$_2$Cl$_2$ and 7·CH$_2$Br$_2$, the guest molecules themselves are oriented more vertically. The ClCH$_2$CH$_2$Cl molecules in 1·ClCH$_2$CH$_2$Cl and 7·ClCH$_2$CH$_2$Cl are also vertically oriented. In 1·CH$_3$I, the CH$_3$I molecule maintains a relatively horizontal position. The distances between the diagonally positioned metal atoms, average distances, and oblateness values, which is defined as 1 – (short radius/long radius), are summarized in Table 1. The M1···M5 and M3···M7 distances correspond to the lengths of the major and minor axes of the ellipsoidal architectures, respectively. In the palladium complexes, the toroidal architecture flexes to accommodate the guest molecules, as is evident from the distances mentioned above (major axis: 8.57–8.79 Å and minor axis: 8.07–8.34 Å), which is similar to the behavior reported for [Ni(μ-SBu)(μ-mtet)]$_{10}$ (mtet = 2-methylthioethanethiolate). In contrast, the platinum complexes did not undergo any significant change in shape in the presence of the guest molecules (major and minor axes lengths of 8.74–8.79 Å and 8.38–8.41 Å, respectively). The oblateness also indicates the difference in the flexibilities of 1 and 7. Among these complexes, the architecture of 1·ClCH$_2$CH$_2$Cl is the closest to a true circle, whereas 1·CH$_2$Cl$_2$ and 1·CH$_2$Br$_2$ have the most ellipsoidal structures.
**Figure 6.** Molecular structures of $1\cdot\text{CH}_2\text{Cl}_2$, $1\cdot\text{ClCH}_2\text{CH}_2\text{Cl}$, $1\cdot\text{CH}_2\text{Br}_2$, and $1\cdot\text{CH}_3\text{I}$ (upper: top views, lower: side views). Methoxycarbonylmethyl groups and solvent molecules outside of the tiara rings are omitted for clarity. The guest molecules are depicted in a space-filling model. Pd: light green, S: orange, Cl: green, Br: brown, I: purple, C: gray, H: white.

**Figure 7.** Molecular structures of $7\cdot\text{CH}_2\text{Cl}_2$, $7\cdot\text{ClCH}_2\text{CH}_2\text{Cl}$, and $7\cdot\text{CH}_2\text{Br}_2$ (upper: top views, lower: side views). Methoxycarbonylmethyl groups and solvent molecules outside of the tiara rings are omitted for clarity. The guest molecules are depicted in a space-filling model. Pt: pink, S: orange, Cl: green, Br: brown, C: gray, H: white.
Table 1. Diagonally Positioned Metal···Metal Distances (Å), Average Distances (Å), and Oblateness Values for 1, 7, and their Inclusion Complexes

<table>
<thead>
<tr>
<th>Complex</th>
<th>M1···M5</th>
<th>M2···M6</th>
<th>M3···M7</th>
<th>M4···M8</th>
<th>Average</th>
<th>Oblateness</th>
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</thead>
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<td>8.2331(9)</td>
<td>8.2307(8)</td>
<td>8.5895(8)</td>
<td>8.44</td>
<td>0.056</td>
</tr>
<tr>
<td>1·CH₂Cl₂</td>
<td>8.7614(9)</td>
<td>8.5509(9)</td>
<td>8.0743(9)</td>
<td>8.4117(9)</td>
<td>8.45</td>
<td>0.078</td>
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<td>1·ClCH₂CH₂Cl</td>
<td>8.5720(19)</td>
<td>8.5711(15)</td>
<td>8.3447(18)</td>
<td>8.3465(16)</td>
<td>8.46</td>
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<tr>
<td>1·CH₂Br₂</td>
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<td>8.4094(13)</td>
<td>8.0890(13)</td>
<td>8.5747(12)</td>
<td>8.46</td>
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<tr>
<td>1·CH₃I</td>
<td>8.7466(6)</td>
<td>8.4364(6)</td>
<td>8.1490(6)</td>
<td>8.5861(6)</td>
<td>8.48</td>
<td>0.068</td>
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<tr>
<td>7</td>
<td>8.8011(5)</td>
<td>8.7539(5)</td>
<td>8.3559(5)</td>
<td>8.4346(5)</td>
<td>8.59</td>
<td>0.051</td>
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<td>7·CH₂Cl₂</td>
<td>8.7880(23)</td>
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<td>7·ClCH₂CH₂Cl</td>
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<td>8.7017(28)</td>
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<tr>
<td>7·CH₂Br₂</td>
<td>8.7935(23)</td>
<td>8.5500(20)</td>
<td>8.4057(20)</td>
<td>8.6355(29)</td>
<td>8.60</td>
<td>0.044</td>
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</table>

As shown in Figure 8, the weak CH···Cl hydrogen bonds between the methylene protons of the four axially positioned methoxycarbonylmethyl groups and the chlorine atom of CH₂Cl₂ in 1·CH₂Cl₂ stabilize the inclusion structure. The H···Cl atomic distances (2.50–2.99 Å) are typical for this type of hydrogen bonding.47-48 Weak CH···X hydrogen bonds (X = Cl, Br) were observed not only in 1·CH₂Cl₂ but also in other dihaloalkane inclusion complexes. The ranges of the corresponding atomic distances and average distances are summarized in Table 2. The shortest distances between the halogen atoms of the guest molecules and palladium atoms in 1·CH₂Cl₂ (3.249(5) Å), 1·CH₂Br₂ (3.187(3) Å), and 1·CH₃I (3.3214(10) Å) are slightly shorter than the sum of the van der Waals radii of palladium (1.63 Å) and the corresponding halogen atoms (Cl: 1.75 Å, Br: 1.85 Å, I: 1.98 Å) by ~0.1, 0.3, and 0.3 Å, respectively; this indicates weak coordination of the halogen atoms to the palladium atoms.
Figure 8. CH···Cl hydrogen bonds in 1·CH₂Cl₂ (top view). Only the tiara ring, four axial methoxycarbonylmethyl groups (two disordered), and a CH₂Cl₂ molecule are shown. The numbers represent the H···Cl atomic distances (Å). Pd: light green, S: orange, Cl: green, C: gray, O: red, H: white.

Table 2. Ranges of CH···X Distances (X = Cl, Br) (Å) and Average Distances (Å) in 1·CH₂Cl₂, 1·ClCH₂CH₂Cl, 1·CH₂Br₂, 7·CH₂Cl₂, 7·ClCH₂CH₂Cl, and 7·CH₂Br₂

<table>
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<tr>
<th>Complex</th>
<th>Range of CH···X Distances</th>
<th>Average</th>
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<tr>
<td>1·CH₂Cl₂</td>
<td>2.50–2.99</td>
<td>2.79(22)</td>
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<tr>
<td>1·ClCH₂CH₂Cl</td>
<td>2.69–3.01</td>
<td>2.87(9)</td>
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<tr>
<td>1·CH₂Br₂</td>
<td>2.47–3.18</td>
<td>2.8(4)</td>
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<tr>
<td>7·CH₂Cl₂</td>
<td>2.55–3.19</td>
<td>2.93(18)</td>
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<tr>
<td>7·ClCH₂CH₂Cl</td>
<td>2.70–2.96</td>
<td>2.83(11)</td>
</tr>
<tr>
<td>7·CH₂Br₂</td>
<td>2.86–3.18</td>
<td>2.98(11)</td>
</tr>
</tbody>
</table>

Conclusion

We synthesized a tiara-like octanuclear palladium complex via a conventional method, which involved a thiol and base, as well as an alternative method using dimethyl [1,3-phenylenebis(1-methylethylidenethio)]diacetate. In the latter method, the resultant tiara-like complex formed via the corresponding SCS–pincer complex and/or 1:1 complexes of PdCl₂ and the ligand. The alternative method is effective for the synthesis of a highly pure tiara-like octanuclear platinum complex. Both of these tiara-like complexes had an ellipsoidal architecture; the platinum tiara ring was more horizontally elongated than the palladium tiara ring because of the differences in the S–M–S and M–S–M bond angles, whereas the M–S bond distances were almost equivalent in the palladium and platinum complexes. In CDCl₃, these complexes showed the dynamic behaviors in which the axial arms would move into and out of the tiara ring, whereas no such behaviors were observed in
CD₂Cl₂. The inclusion of dihalo- or iodoalkanes into the tiara-like complexes was also successful. The toroidal architecture flexed to accommodate the guest molecules in the palladium inclusion complexes, whereas no significant change in the shape occurred upon the inclusion of guest molecules in the platinum inclusion complexes. Weak CH···X hydrogen bonding (X = Cl, Br) as well as weak coordination of the halogen atoms to palladium atoms stabilize these inclusion structures. Other small molecules would also be expected to be accommodated in the inner void of the tiara rings, and some of them may have much stronger interactions with transition-metal atoms resulting in a unique behavior and reactivity.

**Experimental Section**

**Materials and Methods.** All manipulations were performed under an argon atmosphere using standard Schlenk techniques. Dry solvents were purchased from either Wako Chemical or Nacalai. Flash column chromatography was performed using silica gel SILICYCLE SiliaFlash F60 (40-63 μm, 230-400 mesh).

**Physical and Analytical Measurements.** NMR spectra were recorded on either a JEOL AL-400 (400 MHz (¹H), 100 MHz (¹³C), 85.7 MHz (¹⁹⁵Pt)) or a Bruker AV-300N (300 MHz (¹H), 75 MHz (¹³C)) spectrometer. Chemical shift values (δ) in ¹H and ¹³C NMR spectra were expressed relative to SiMe₄. ¹⁹⁵Pt NMR spectra referenced using an external reference (0.3 M K₂PtCl₄ in D₂O, δ = -1628 ppm). IR measurements were carried out using a JASCO FT/IR-6100 spectrometer. Elemental analysis was obtained using a J-SCIENCE LAB JM-10 analyzer. MALDI-TOF MS spectra were recorded on a SHIMADZU AXIMA Performance spectrometer. Melting points were measured on a Yanagimoto micro melting point apparatus.

**Synthesis of [Pd(µ-SCH₂CO₂Me)₂]₈ (1).** To a mixture of methyl thioglycolate (51 μL, 0.57 mmol), DIEA (96 μL, 0.56 mmol), and n-propanol (0.95 mL), PdCl₂ (50 mg, 0.28 mmol) was added and the reaction mixture was stirred at room temperature for 4 h. After filtration, the orange-yellow solid was dried under vacuum and obtained in 90% isolated yield (81 mg, 0.032 mmol). Recrystallization from CHCl₃/hexane gave yellow block crystals. ¹H NMR (400 MHz, CDCl₃) δ 3.97 (s, 24H, ax-CO₂CH₃), 3.77 (s, 24H, eq-CO₂CH₃), 3.34 (s, 16H, ax-CH₂), 3.23 (s, 16H, eq-CH₂). ¹H NMR (400 MHz, CDCl₃) δ 3.75 (s, 24H, CO₂CH₃), 3.72 (s, 24H, CO₂CH₃), 3.38 (s, 16H, CH₂), 3.22 (s, 16H, CH₂). ¹³C NMR (75 MHz, CDCl₃) δ 170.4 (ax-CO₂Me), 169.8 (eq-CO₂Me), 53.1 (ax-CO₂CH₃), 52.7 (eq-CO₂CH₃), 35.6 (eq-CH₂), 30.6 (ax-CH₂). IR (νCO, KBr): 1735 cm⁻¹. Mp 143–144 °C (dec.). Anal. Calcd for C₄₈H₈₀O₃₂Pd₈S₁₆: C, 22.76; H, 3.18. Found: C, 22.65; H, 3.02.

**Synthesis of Dimethyl [1,3-Phenylenebis(1-methylethylidenethio)]diacetate (2).** Compound 2 was prepared by reference to a reported synthetic procedure for similar compounds.⁴⁹ To a mixture of α,α'-dihydroxy-1,3-diisopropylbenzene (776 mg, 3.99 mmol), ZnI₂ (1.28 g, 4.01 mmol), and methyl
thioglycolate (0.750 mL, 8.39 mmol), dry 1,2-dichloroethane (25 mL) was added and the reaction mixture was stirred at room temperature for 1 h. The mixture was diluted with dichloromethane and washed with water. The organic layer was then washed with 1 M NaOH and dried over MgSO₄. After filtration and removal of the solvents, the residue was purified by silica gel column chromatography (eluent: hexane/ethyl acetate = 3:1) to give 2 (1.42 g, 3.82 mmol) in 96% isolated yield as a colorless oil. ^1H NMR (300 MHz, CDCl₃) δ 7.71 (t, J = 1.5 Hz, 1H, Ar-2-H), 7.37 (dd, J = 6.6, 1.5 Hz, 2H, Ar-4,6-H), 7.28 (t, J = 6.9 Hz, 1H, Ar-5-H), 3.59 (s, 6H, CO₂CH₃), 2.98 (s, 4H, CH₂), 1.73 (s, 12H, CH₃). ^13C NMR (75 MHz, CDCl₃) δ 171.0 (CO₂Me), 145.2 (Ar-1,3), 127.9 (Ar-5), 125.1 (Ar-4,6), 125.0 (Ar-2), 52.3 (CO₂CH₃), 48.6 (C(CH₃)₂), 32.1 (SCH₂), 29.9 (C(CH₃)₂).

Alternative Synthesis of 1. To a mixture of [PdCl₂(MeCN)₂] (78 mg, 0.30 mmol) and 2 (112 mg, 0.30 mmol), dry methanol (1.0 mL) was added and the reaction mixture was heated at 70 °C (bath temperature) for 10 h. After filtration, the orange-yellow precipitate was dried under vacuum and obtained in 84% isolated yield (80 mg, 0.031 mmol). Recrystallization from CHCl₃/hexane gave yellow block crystals.

Synthesis of SCS-Pincer Palladium Complex 3. To a mixture of [PdCl₂(MeCN)₂] (1.00 g, 3.86 mmol) and 2 (1.43 g, 3.86 mmol), dry chloroform (100 mL) was added and the reaction mixture was stirred at room temperature for 1 h. The obtained crude compound was purified by silica gel column chromatography (eluent: hexane/ethyl acetate = 1:2) to give 3 as a pale orange-yellow solid (1.64 g, 3.20 mmol) in 83% isolated yield. ^1H NMR (300 MHz, CDCl₃) δ 7.11 (t, J = 7.8 Hz, 1H, Ar-4-H), 6.80 (d, J = 7.5 Hz, 2H, Ar-3,5-H), 4.50–3.85 (br, 4H, CH₂), 3.76 (s, 6H, CO₂CH₃), 1.81 (s, 12H, CH₃). ^13C NMR (100 MHz, CDCl₃) δ 167.1 (CO₂Me), 154.1 (Ar-2,6), 152.7 (br, Ar-1), 125.0 (Ar-4), 123.1 (Ar-3,5), 63.4 (C(CH₃)₂), 52.3 (CO₂CH₃), 35.7 (SCH₂), 31.1 (br, CH₃), 29.4 (br, CH₃). IR (νCO, KBr): 1739 cm⁻¹. Mp 119–121 °C (dec.). Anal. Calcd for C₁₈H₂₅ClO₄PdS₂: C, 42.28; H, 4.93. Found: C, 42.31; H, 4.85.

Synthesis of Dimethyl [1,3-Phenylenebis(methylene)diacetate] (5). To a mixture of methyl thioglycolate (240 μL, 2.68 mmol) and DIEA (455 μL, 2.68 mmol), dry methanol (10 mL) was added and the reaction mixture was stirred at room temperature for 15 min. α,α'-Dibromo-m-xylene (335 mg, 1.27 mmol) was then added and the mixture was stirred at 70 °C for 10 h. After drying under vacuum, the obtained crude compound was purified by silica gel column chromatography (eluent: hexane/ethyl acetate = 3:1) to give 5 (392 mg, 1.25 mmol) in 98% isolated yield as a colorless oil. ^1H NMR (300 MHz, CDCl₃) δ 7.31–7.21 (m, 4H, Ar-2,4,5,6-H), 3.81 (s, 4H, ArCH₂), 3.72 (s, 6H, CO₂CH₃), 3.08 (s, 4H, CH₂CO₂Me). ^13C NMR (75 MHz, CDCl₃) δ 170.5 (CO₂Me), 137.3 (Ar-1,3), 129.6 (Ar-5), 128.5 (Ar-2), 127.9 (Ar-4,6), 52.1 (CO₂CH₃), 35.9 (ArCH₂), 31.8 (CH₂CO₂Me).
Synthesis of SCS-Pincer Palladium Complex 6. To a mixture of [PdCl₂(MeCN)]₂ (15 mg, 0.059 mmol) and 5 (19 mg, 0.060 mmol), dry acetonitrile (1.5 mL) was added and the reaction mixture was stirred at 90 °C (bath temperature) for 20 h. The yellow solution was vacuumed to dry. Recrystallization from CHCl₃/hexane gave a pale yellow solid (15 mg, 0.033 mmol) in 57% isolated yield. Recrystallization from CHCl₃/hexane gave pale yellow block crystals. ¹H NMR (300 MHz, CDCl₃) δ 7.04–6.94 (m, 3H, Ar-3,4,5-H), 4.60–3.90 (br, 8H, CH₂), 3.79 (s, 6H, CO₂CH₃). ¹³C NMR (100 MHz, CDCl₃) δ 167.6 (CO₂Me), 158.5 (br, Ar-1), 147.6 (br, Ar-2,6), 125.2 (Ar-4), 122.7 (Ar-3,5), 53.1 (CO₂CH₃), 48.5 (br, ArCH₂), 39.9 (br, CH₂CO₂Me). IR (ν₁CO, KBr): 1734 cm⁻¹. Mp 112–113 °C (dec.). Anal. Calcd for C₁₄H₁₇ClO₄PdS₂: C, 36.93; H, 3.76. Found: C, 36.58; H, 3.99.

Synthesis of [Pt(µ-SCH₂CO₂Me)₂]₈ (7). To a mixture of PtCl₂ (662 mg, 2.49 mmol) and 2 (922 mg, 2.49 mmol), dry acetonitrile (1.25 mL) was added and the reaction mixture was stirred at 90 °C (bath temperature) for 10 h. The yellow solution was evaporated under vacuum to dry. Recrystallization from CHCl₃/heptane gave yellow block crystals (950 mg, 0.293 mmol) in 94% isolated yield. ¹H NMR (400 MHz, CDCl₃) δ 3.96 (s, 24H, ax-CO₂CH₃), 3.76 (s, 24H, eq-CO₂CH₃), 3.53 (s, 16H, ax-CH₂), 3.42 (s, 16H, eq-CH₂). ¹H NMR (400 MHz, CD₂Cl₂) δ 3.75 (s, 24H, CO₂CH₃), 3.70 (s, 24H, CO₂CH₃), 3.60 (s, 16H, CH₂), 3.40 (s, 16H, CH₂). ¹³C NMR (75 MHz, CDCl₃) δ 170.1 (eq-CO₂Me), 169.9 (ax-CO₂Me), 52.9 (ax-CO₂CH₃), 52.6 (eq-CO₂CH₃), 37.1 (eq-CH₂), 30.2 (ax-CH₂). ¹⁹⁵Pt NMR (85.7 MHz, CDCl₃) δ -3247.3 (br). IR (ν₁CO, KBr): 1733 cm⁻¹. Mp 178–180 °C (dec.). Anal. Calcd for C₄₈H₈₀O₃₂Pt₈S₈: C, 17.78; H, 2.49. Found: C, 18.11; H, 2.55.

SCS-Pincer Platinum Complex 8. To a mixture of PtCl₂ (369 mg, 1.39 mmol) and 2 (513 mg, 1.39 mmol), dry acetonitrile (36 mL) was added and the reaction mixture was stirred at 90 °C (bath temperature) for 20 h. The yellow solution was evaporated under vacuum to dry. The resulting yellow oil was dissolved in CHCl₃, and hexane was poured into the solution to remove 7. After filtration, the filtrate was dried under vacuum, and the obtained crude compound was purified by silica gel column chromatography (eluent: hexane/ethyl acetate = 1:1) to give 8 (124 mg, 0.206 mmol) in 15% isolated yield as a pale yellow solid. ¹H NMR (300 MHz, CD₃CN) δ 7.13 (t, J = 7.8 Hz, 1H, Ar-4-H), 6.86 (d, J = 7.5 Hz, 2H, Ar-3,5-H), 4.08–3.73 (m, 4H, CH₂), 3.67 (s, 6H, CO₂CH₃), 1.75 (s, 12H, CH₃). ¹³C NMR (100 MHz, CD₃CN) δ 168.3 (t, 3J_C-Pt = 23.8 Hz, CO₂Me), 153.8 (t, 2J_C-Pt = 28.8 Hz, Ar-2,6), 145.1 (Ar-1), 144.6 (Ar-1), 125.4 (Ar-4), 123.8 (t, 3J_C-Pt = 23.2 Hz, Ar-3,5), 67.8 (t, 2J_C-Pt = 18.7 Hz, C(CH₃)₂), 53.4 (CO₂CH₃), 38.6 (SCH₂), 37.9 (SCH₂), 32.0 (CH₃), 31.3 (CH₃), 29.4 (CH₃), 28.6 (CH₃). ¹⁹⁵Pt NMR (85.7 MHz, CD₃CN) δ -4026.5 (quintet, 3J_Pt-H = 42 Hz), -4028.2 (quintet, 3J_Pt-H = 38 Hz). IR (ν₁CO, KBr): 1736 cm⁻¹. Mp 131.5–132.5 °C (dec.). Anal. Calcd for C₁₈H₂₅ClO₄PdS₂: C, 36.03; H, 4.20. Found: C, 36.11; H, 4.11.
Crystallographic Study of Complexes 1, 6, 7, 1·CH₂Cl₂, 1·ClCH₂CH₂Cl, 1·CH₂Br₂, 1·CH₃I, 7·CH₂Cl₂, 7·ClCH₂CH₂Cl, and 7·CH₂Br₂. Crystals suitable for X-ray diffraction measurements were obtained by recrystallization from CHCl₃/hexane for 1 (yellow block crystals), CH₂Cl₂/hexane for 6 (pale yellow block crystals), CHCl₃/octane for 7 (yellow block crystals), CH₂Cl₂/Et₂O for 1·CH₂Cl₂ (yellow prism crystals), ClCH₂CH₂Cl/hexane for 1·ClCH₂CH₂Cl (yellow rod crystals), CH₂Br₂/hexane for 1·CH₂Br₂ (yellow rod crystals), CH₃I/Et₂O for 1·CH₃I (yellow block crystals), CH₂Cl₂/hexane for 7·CH₂Cl₂ (yellow block crystals), ClCH₂CH₂Cl/Et₂O for 7·ClCH₂CH₂Cl (yellow block crystals), and CH₂Br₂/hexane for 7·CH₂Br₂ (yellow cube crystals), respectively. The crystals dipped in liquid paraffin were mounted using a cryoloop and frozen at either 153 or 123 K. The diffraction data were collected with a Rigaku Saturn CCD detector (MoKα, λ = 0.71073 Å). Crystal data and structure refinement parameters are listed in Tables S1–S3 in Supporting Information. The structures were solved by direct methods using either SHELXS-201350 (for 1, 1·CH₂Cl₂, 1·ClCH₂CH₂Cl, 1·CH₂Br₂, and 1·CH₃I), SHELXS-9751 (for 6), or SIR9752 (for 7, 7·CH₂Cl₂, 7·ClCH₂CH₂Cl, and 7·CH₂Br₂) and refined by least squares on F², SHELXL-2013 or SHELXL-97 (for 6).50-53 Non-hydrogen atoms were anisotropically refined except for the disordered ones. Because methoxycarbonylmethyl groups and small organic molecules were disordered occasionally, bond distances in these moieties were restrained to be equal, respectively, using a SADI command in SHELXL. All hydrogen atoms were placed at calculated positions. Refinements were continued until all shifts were smaller than one-tenth of the standard deviations of the parameters involved. Atomic scattering factors and anomalous dispersion terms were taken from the International Tables for X-ray Crystallography.54

MALDI-TOF MS Analysis of Tiara-like Complexes 1 and 7. Acetonitrile solutions of the tiara-like complexes (0.5 μL, 1: 9.9 mM, 7: 7.9 mM) were mixed with a H₂O/ethanol (1:1) solution of all-trans-retinoic acid (0.5 μL, 50.0 mM) and TFA aq. (0.5 μL, 0.1%) on a MALDI plate and were analyzed. Calibration was performed using bradykinin fragment 1–7 in 0.1% TFA aq. (10 μM), adrenocorticotropic hormone (ACTH) fragment 18-39 (human) in 0.1% TFA aq. (10 μM), and α-cyano-4-hydroxycinnamic acid (4-CHCA).

Associated Content
Supporting Information. NMR spectra for new compounds, MALDI-TOF MS spectra for 1 and 7, crystal data and structure refinement parameters for 1, 6, 7, 1·CH₂Cl₂, 1·ClCH₂CH₂Cl, 1·CH₂Br₂, 1·CH₃I, 7·CH₂Cl₂, 7·ClCH₂CH₂Cl, and 7·CH₂Br₂ and CIF files for 1, 6, 7, 1·CH₂Cl₂, 1·ClCH₂CH₂Cl, 1·CH₂Br₂, 1·CH₃I, 7·CH₂Cl₂, 7·ClCH₂CH₂Cl, and 7·CH₂Br₂. CCDC reference numbers 974025 (1), 974026 (6), 974027 (7), 974058 (1·CH₂Cl₂), 974059 (1·ClCH₂CH₂Cl), 974060 (1·CH₂Br₂), 976482 (1·CH₃I), 976483 (7·CH₂Cl₂), 976484 (7·ClCH₂CH₂Cl), and 976485 (7·CH₂Br₂). This material is available free of charge via the Internet at http://pubs.acs.org.
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Notes. The authors declare no competing financial interest.

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References
Table of Contents Synopsis
A tiara-like octanuclear palladium thiolate complex, $[\text{Pd}(\mu-\text{SCH}_2\text{CO}_2\text{Me})_2]_8$, was synthesized via reactions of either $\text{PdCl}_2$ with methyl thioglycolate/$\text{N}_2\text{N}$-diisopropylethylamine or $[\text{PdCl}_2(\text{MeCN})_2]$ with $m$-$\text{C}_6\text{H}_4(\text{CMe}_2\text{SCH}_2\text{CO}_2\text{Me})_2$. The latter method efficiently produced the corresponding platinum complex in high purity. Small molecules such as $\text{CH}_2\text{Cl}_2$ were accommodated in the inner voids. These complexes are stabilized by weak CH···X hydrogen bonds ($X = \text{Cl}$ or $\text{Br}$), as well as weak coordination of the halogen atoms to the transition-metal atoms.

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